

# Atomic Theory Development

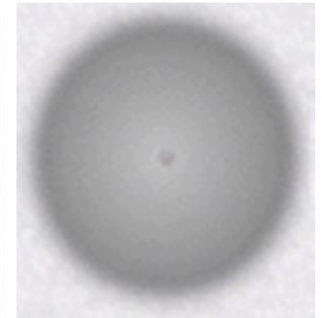
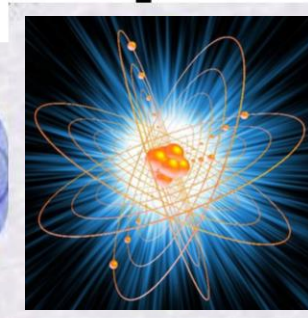
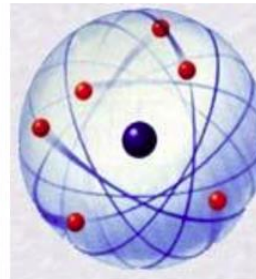
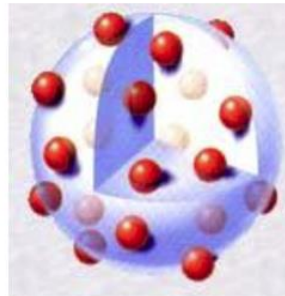
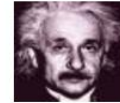
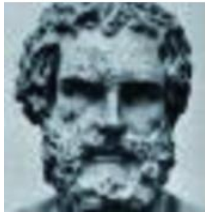
Democritus 460 BC  
and Dalton 1803 AD

Thomson  
1897

Rutherford  
1912

Bohr  
1913

Modern  
Quantum  
Cloud Model  
post 1930



Born **as early as 400 BC**, it took more than 2000 years before Science was ready to accept the idea of atomic structure of matter...and another 150 years to develop a good model!

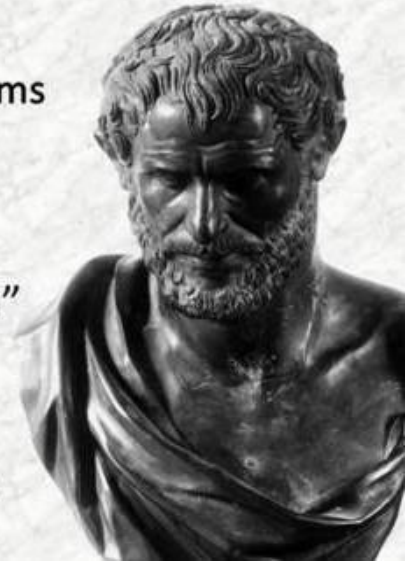
# Democritus

## ~400 BC

“atomos”=“not to be cut”

“Nothing exists except atoms  
and empty space;  
everything else is opinion”

*Democritus*  
(ca. 460 BC – ca. 370 BC)



- Matter **could not** be divided into smaller and smaller pieces forever, eventually the **smallest possible piece** would be obtained.
- This piece, **atomos** (atom), would be **indivisible**.
- Between atoms, there would be **empty space**.
- To Democritus, atoms were **small, hard particles of different shapes and sizes** that were **all made of the same material**.
- Atoms were infinite in number, always moving and capable of joining together.

# John Dalton

## early 1800s



The **first truly scientific theory of the atom**: conclusions were reached by experimentation and examination of the results in an empirical fashion.

- All **elements** are composed of **atoms**.
- Atoms are indivisible and indestructible particles.
- Atom model: a billiard ball or a *marble*.

- H**  
**O**  
**W**  
**?**
- Atoms of the same element are exactly alike.
  - Atoms of different elements are different.
  - Compounds are formed by the joining of atoms of two or more elements.



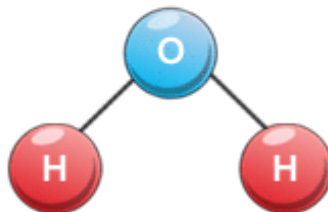
**Color?**  
**Size?**

**Mass?**

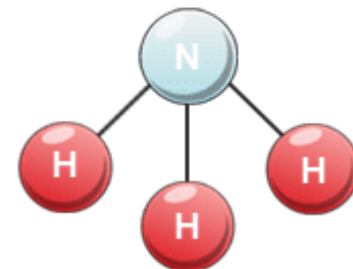
# How to find Atomic Weight?

- Assumption #1: atoms of a given chemical element are identical in size, **mass**, and all other properties.
- Assumption #2: chemical **compounds** are formed when atoms of different elements combine in simple whole-number ratios.

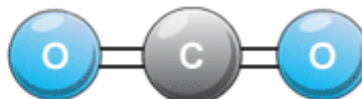
WATER



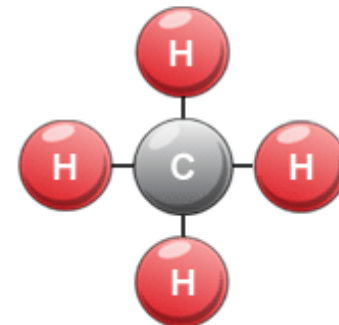
AMMONIA



CARBON  
DIOXIDE



METHANE



Atomic weights of elements can be determined by careful **weighing** of chemical reactions!

## John Dalton

- 1803-1805: **first list** of **relative atomic weights** containing just **6 elements**, namely *hydrogen* (conventionally assumed to weigh 1), *oxygen*, *nitrogen*, *carbon*, *sulfur*, and *phosphorus*.

- 1808: expanded list of elements

ELEMENTS			
○	Hydrogen 1	⊕	Strontian 46
⊖	Nitrogen 5	⊗	Barytes 68
●	Carbon 5	⊙	Iron 50
○	Oxygen 7	⊕	Zinc 56
⊖	Phosphorus 9	⊗	Copper 56
⊕	Sulphur 13	⊙	Lead 90
⊖	Magnesia 20	⊕	Silver 190
⊕	Lime 24	⊗	Gold 190
⊖	Soda 28	⊙	Platina 190
⊕	Potash 42	⊗	Mercury 167

## Dmitri Mendeleev

- 1869: **periodic table** of **66 elements** ordered and grouped according to their **atomic weight**.

ОПЫТЪ СИСТЕМЫ ЭЛЕМЕНТОВЪ.					
ОСНОВАННОЙ НА ВѢСЪ АТОМНОМЪ ВѢСѢ И ХИМИЧЕСКОМЪ СХОДСТВѢ.					
	Ti = 50	Zr = 90	? = 180.		
	V = 51	Nb = 94	Ta = 182.		
	Cr = 52	Mo = 96	W = 186.		
	Mn = 55	Rh = 104,4	Pt = 197,4		
	Fe = 56	Ru = 104,4	Ir = 198.		
	Ni = 59	Pd = 106,4	Os = 199.		
H = 1	Cu = 63,4	Ag = 108	Hg = 200.		
Be = 9,4	Mg = 24	Zn = 65,4	Cd = 112		
B = 11	Al = 27,4	? = 68	U = 116	Au = 197,7	
C = 12	Si = 28	? = 70	Sn = 118		
N = 14	P = 31	As = 75	Sb = 122	Bi = 210?	
O = 16	S = 32	Se = 79,4	Te = 128?		
F = 19	Cl = 35,4	Br = 80	I = 127		
Li = 7	Na = 23	K = 39	Rb = 85,4	Cs = 133	Tl = 204.
		Ca = 40	Sr = 87,4	Ba = 137	Pb = 207.
		? = 45	Ce = 92		
		?Er = 56	La = 94		
		?Yt = 60	Di = 95		
		?In = 75,4	Th = 118?		

Д. Менделѣевъ

# Scientific Mysteries of 1870s

## LIGHTEST

1 <b>H</b> Hydrogen 1	Atomic # Symbol Name Atomic weight
3 <b>Li</b> Lithium 7	4 <b>Be</b> Beryllium 9
11 <b>Na</b> Sodium 23	12 <b>Mg</b> Magnesium 24

Elements are grouped and *ordered* according to their **atomic weight**...

## Fragment of the Periodic Table (showing elements known by 1869 when Mendeleev published his first version)

5 <b>B</b> Boron 11	6 <b>C</b> Carbon 12	7 <b>N</b> Nitrogen 14	8 <b>O</b> Oxygen 16	9 <b>F</b> Fluorine 19	10 <b>Ne</b> Neon 20.1797
13 <b>Al</b> Aluminium 27	14 <b>Si</b> Silicon 28	15 <b>P</b> Phosphorus 31	16 <b>S</b> Sulfur 32	17 <b>Cl</b> Chlorine 35	18 <b>Ar</b> Argon 39.948
31 <b>Ga</b> Gallium 69.723	32 <b>Ge</b> Germanium 72.64	33 <b>As</b> Arsenic 75	34 <b>Se</b> Selenium 79	35 <b>Br</b> Bromine 80	36 <b>Kr</b> Krypton 83.795
49 <b>In</b> Indium 115	50 <b>Sn</b> Tin 118	51 <b>Sb</b> Antimony 122	52 <b>Te</b> Tellurium 128	53 <b>I</b> Iodine 127	54 <b>Xe</b> Xenon 131.293

UNKNOWN

...but not always!

**Puzzling question: what carries electricity?**