

HW28, May 8th.

Bases

Let's remember how should we write structural formula and Lewis structure for ionic compounds. Sodium chloride, chemical formula NaCl, structural formula Na^+Cl^- .

When we write down Lewis structures for ionic compounds, you should remember the following: there are atoms that give away electron(s), and there are atoms that will accept electron(s), and we indicate corresponding charges on the atoms.

Sodium chloride formation looks like this:



Bases, metal hydroxides

Composed of metal atom(s) and hydroxyl group(s) e.g. NaOH, $\text{Ca}(\text{OH})_2$

1. React with acidic oxides forming salt and water.
2. React with acids forming salt and water.
3. Can decompose upon heating into basic oxide and water.

Questions:

1. Write chemical equations for the following transformations:



2. Write the structural formula and Lewis structure for NaOH and $\text{Mg}(\text{OH})_2$