SchoolNova Chem 2 Class 17 Homework - Enthalpy, Entropy, Free Energy

1. For an exothermic reaction, what must be true about the sign of ΔH (enthalpy)?

2. For the reaction $A(g) \leftrightarrow B(l)$, do you predict that there will be a positive or negative value for ΔS (entropy) in the forward direction?

3. When relating the equilibrium constant (Keq) to ΔG (Gibbs free energy), what measurable variable must also be considered?