

Name: _____

Periodic Table Quiz 2

Identify the element and characterize as Metal, Transition Metal, Metalloid, Non Metal or Noble Gas.

- 1) Period 6, Group 11 _____
- 2) Period 3 Group 17 _____
- 3) Period 2 Group 18 _____
- 4) Period 5 Group 2 _____
- 5) Period 4 Group 1 _____

6) Describe the element that occupies the position of Period 3 Group 14

- a) Symbol _____
- b) Name _____
- c) Characteristics _____
- d) Valence Electrons _____
- e) Atomic Number _____
- f) Atomic Mass _____
- g) Electron Configuration _____
- h) Number of Electrons _____
- i) Number of Protons _____
- j) Charge of Nucleus _____
- k) Number of Energy Levels or Shells _____

7) Why are elements in the same group? What do elements in the same group have in common?

8) The horizontal rows on the periodic table are called periods. What do elements in the same period have in common?

9) Does the atomic radius increase or decrease as you travel down a group? Why?

10) Does the atomic radius of atoms increase or decrease as you travel across a period? Why?

11) In ground state, atoms of the elements in Group 15 of the periodic table all have the same number of

- a) filled principal energy levels b) occupied principal energy levels
- c) neutrons in the nucleus d) electrons in the valence levels

12) The properties of Carbon are expected to be most similar to those of

- a) Boron b) Aluminum c) Silicon d) Phosphorus

13) An atom has 28 innermost electrons and 7 valence electrons. In which period of the periodic table is this element located?

- a) 5 b) 2 c) 3 d) 4

16) Aqueous solutions of compounds containing element X are blue. Element X could be

- a) Carbon b) Copper c) Sodium d) Potassium

17) Which set of elements contains a metalloid?

- a) Fr, F, O, Rn b) K, As, Ar, Mn c) Ba, Ag, Sn, Xe d) Li, Mg, Ca, Kr

19) The element Chlorine (Cl) has properties of

- a) metals only b) metals and nonmetals c) nonmetals only d) neither metals or nonmetals