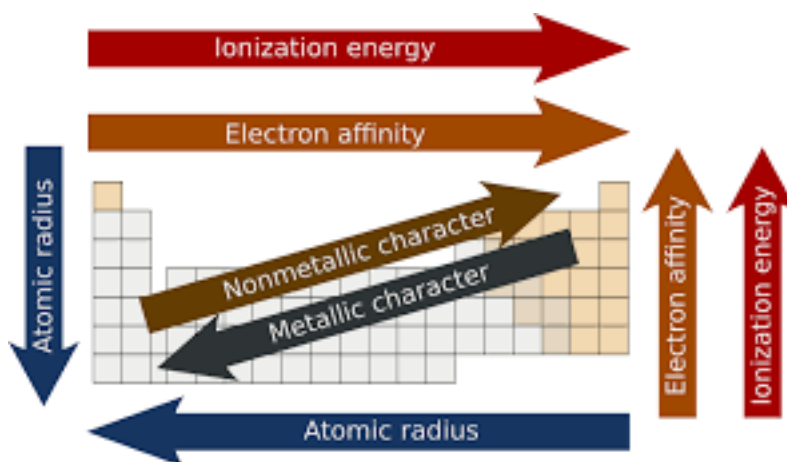


Periodic Trends Worksheet

You can use Table S to figure this out and the table below



- 1) Rank the following elements by increasing atomic radius:
carbon, aluminum, oxygen, potassium.

- 2) Rank the following elements by increasing electronegativity:
sulfur, oxygen, neon, aluminum.

- 3) What is the difference between electron affinity and ionization energy?

- 4) Why does fluorine have a higher ionization energy than iodine?

5) Why do elements in the same family generally have similar properties?

6. In each of the following pairs, circle the species with the *larger* atomic radius:

(a) Mg or Ba

(b) S or S²⁻

(c) Cu⁺² or Cu

(d) He or H⁻

(e) Na or Cl