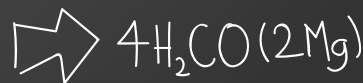
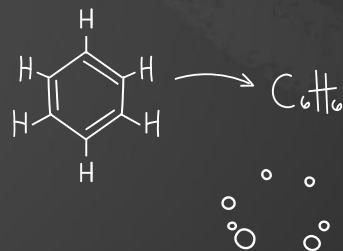
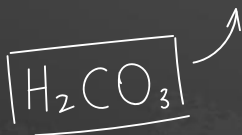
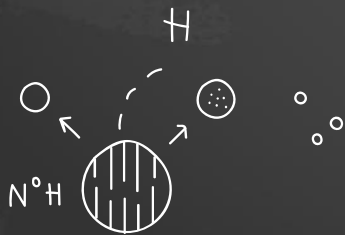
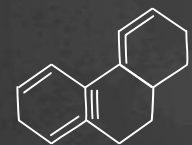




Chemistry - 101

January 23



Moles - unit of counting used in chemistry

Avogadro's number (6.022×10^{23}) represents the number of particles (atoms, ions, formula units, or molecules) in one mole of any substance

Any atom, element, or compound can have its mass expressed in atomic mass units (amu). The average atomic mass (in amu) for atoms of any element can be found on the periodic table.

Mole and amu

The numerical value of the amu of the atoms of a given element is equal to the mass in grams of one mole of that element.

The mass of a single Na atom is 22.99 amu
One mole of Na has an average mass of 22.99 grams

Conversions of # of particles, # of moles, and the mass

E.g. NaCl has a mass of $(22.99 + 35.45) = 58.44$ amu

One mole of NaCl has a mass of $(22.99 + 35.45) = 58.44$ g

This class uses the materials from the following books:

Larry Gonick and Graig Criddle “The cartoon guide to chemistry”

Manyuilov and Rodionov “Chemistry for children and adults”

Kuzmenko, Eremin, Popkov “Beginnings of chemistry”