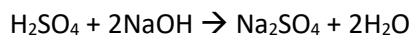
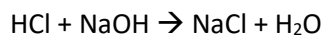
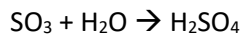


**Classes of chemical compounds - 2**

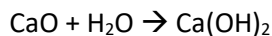
- A. Reactions where acids and bases react with each other are called **reactions of neutralization**. In these reactions a salt and water are formed. E.g. below is a neutralization reaction between hydrochloric acid (HCl – acid) and sodium hydroxide (NaOH – base) with formation of salt (sodium chloride, NaCl) and water:



- B. When acidic oxides react with water, they form acids. E.g.:

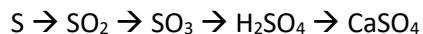
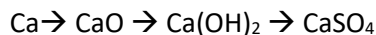


- C. When basic oxides react with water, they form bases. E.g.:



1. Using Periodic Table of Elements write chemical formulas of oxides for the following elements: K, Ba, Fe(II), Cr(III), Cl(VII), Si(IV). Underline formulas of acidic oxides.

2. Write chemical equations for the following transformations:



3. How many grams of concentrated sulfuric acid (H<sub>2</sub>SO<sub>4</sub>) will be necessary to neutralize water solution containing 4 g of NaOH? Assume 100% sulfuric acid concentration.