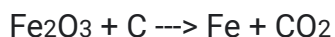


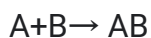
**Chemistry 0 Unit 2 Review Test**  
**Assigned on 02/21/2021    Due date: 02/27/2021**

1. Freezing is a physical change.  
True or false: \_\_\_\_\_.
2. Are fireworks chemical change or physical change?
3. An example of a chemical change is: \_\_\_\_\_.
  - A. chocolate syrup mixed in milk
  - B. breaking glass
  - C. melting ice
  - D. burning coal
4. What is the naming for LiBr? \_\_\_\_\_.
  - A. Lithium(II) Bromide
  - B. Lithium Bromide
  - C. Lithium Boron
  - D. Lithium(II) Boron
5. What is the name of the compound with the formula  $\text{PCl}_3$ ? \_\_\_\_\_.
  - A. Monopotassium Trichloride
  - B. Monophosphorus Trichloride
  - C. Phosphorus Trichloride
  - D. Potassium Trichloride
6. Number of total atoms in  $\text{Ca}(\text{OH})_2$ : \_\_\_\_\_.
7. Number of total atoms in  $\text{Al}_2(\text{Cr}_2\text{O}_7)_3$ : \_\_\_\_\_.
8. What is the correct balancing for this equation?\_\_\_\_\_.
  - A.  $2\text{CH}_4 + 4\text{O}_2 \rightarrow 2\text{CO}_2 + 5\text{H}_2\text{O}$
  - B.  $5\text{CH}_4 + 3\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
  - C.  $\text{CH}_4 + 2\text{O}_2 \rightarrow \text{CO}_2 + 2\text{H}_2\text{O}$
  - D.  $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
9. Please balance the following chemical equation:  
 $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$

10. Please balance the following chemical equation:

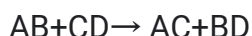


11. What kind of chemical reaction is this?



- A. Double Replacement
- B. Synthesis
- C. Decomposition
- D. Combustion

12. What kind of chemical reaction is this?

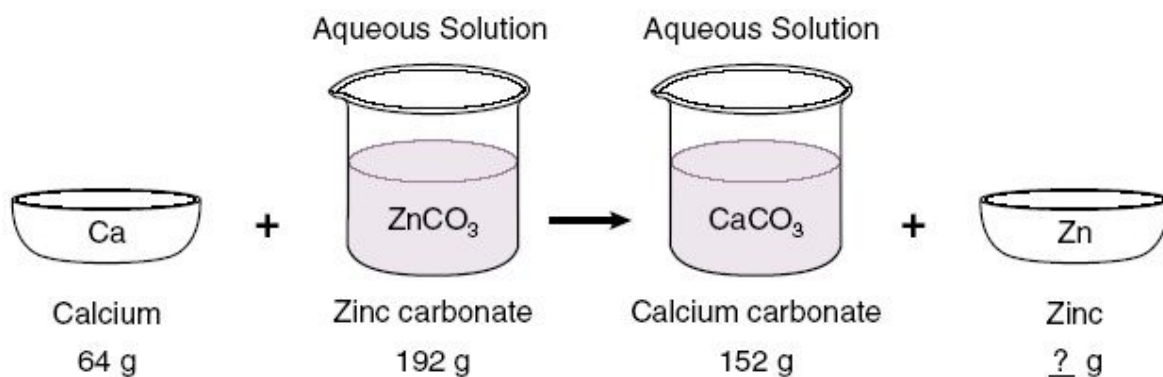


- A. Double Replacement
- B. Synthesis
- C. Decomposition
- D. Combustion

13. The Law of Conservation of Mass states:

- A. Matter can be created and destroyed.
- B. Matter can not be created but it can be destroyed.
- C. Matter cannot be created and it cannot be destroyed.
- D. Matter is not real.

14. Find the missing mass: \_\_\_\_\_.



- A. 104 g
- B. 256 g
- C. 40 g
- D. 88 g

15. If the temperature of a reaction is decreased, what effect will it have on the rate of reaction? \_\_\_\_\_.
- A. It will have no effect.
  - B. The reaction will stop.
  - C. The rate of reaction will decrease.
  - D. The rate of reaction will increase.
16. During the glow stick experiment, what did we keep as our control?\_\_\_\_\_.
- A. The glow stick in hot water.
  - B. The glow stick in the dry ice.
  - C. The glow stick at room temperature.
  - D. There was no control.
17. Why does a catalyst increase the rate of reaction?
- A. It increases the activation energy the particles have.
  - B. It makes more collisions happen.
  - C. It adds more particles.
  - D. It provides an alternative route of lower activation energy.
18. A catalyst \_\_\_\_\_ the energy needed for a chemical reaction and \_\_\_\_\_ the rate of the reaction.
- A. decreases, increases
  - B. increases, decreases
  - C. balances, increases
  - D. balances, decreases
19. Endothermic or exothermic?\_\_\_\_\_.



- A. Endothermic
- B. Exothermic

20. An endothermic reaction feels\_\_\_\_\_.
- A. hot
  - B. cold
  - C. neither hot nor cold
21. When an acid and a base are mixed they\_\_\_\_\_.
- A. explode
  - B. dissolve
  - C. bubble
  - D. neutralize each other
22. Is sodium hydroxide an acid or base?
- A. Acid
  - B. Base
23. A pH of 3 is less acidic than a pH of 5?
- A. True
  - B. False
24. What is the end point of a titration?
- A. When an indicator changes color
  - B. When a acid fully ionizes
  - C. when a base fully ionizes
  - D. when an indicator is added to a solution
25. Define molarity: \_\_\_\_\_.
- A. a substance that changes color inside of an acid or base
  - B. a substance whose particles are dissolved in a solution
  - C. the # of moles of a dissolved solute per liter of solution
  - D. none of the above