

Chemistry 0 Week 17 HW
Assigned on 02/07/2021 Due date:02/13/2021

1. You are titrating a base into an acid to determine the concentration of the base. The endpoint of the neutralization is reached but the stopcock on the Buret sticks slightly and allows a few more drops of base to fall into the solution. How will this affect your calculations for the concentration of the acid?

2. If it takes 21 grams of baking soda to neutralize a beaker of acetic acid, how many moles of acetic acid do you have? If there is 1 L of the acetic acid solution in the beaker, what is the concentration of the acetic acid?

3. It takes 80 mL of a 0.5 M NaOH solution to neutralize 200 mL of an HCl solution. What is the concentration of the HCl solution?

4. Can I titrate a solution of unknown concentration with another solution with unknown concentration and still get a meaningful answer? Please explain.