








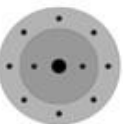
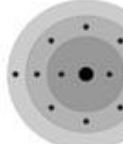






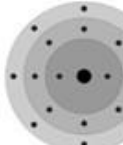


<b>ENERGY LEVELS ELEMENTS 1-20</b>		Complete each energy level model by drawing the correct number of electrons in their corresponding energy levels.															
<b>HYDROGEN</b> 1  1.01	<b>HELIUM</b> 2  4.00	<b>LITHIUM</b> 3  6.94	<b>BERYLLIUM</b> 4  9.01	<b>BORON</b> 5  10.81	<b>CARBON</b> 6  12.01	<b>NITROGEN</b> 7  14.01	<b>OXYGEN</b> 8  16.00	<b>FLUORINE</b> 9  19.00	<b>NEON</b> 10  20.18	<b>SODIUM</b> 11  22.99	<b>MAGNESIUM</b> 12  24.31	<b>ALUMINUM</b> 13  26.98	<b>SILICON</b> 14  28.09	<b>PHOSPHORUS</b> 15  30.97	<b>SULFUR</b> 16  32.07	<b>CHLORINE</b> 17  35.45	<b>ARGON</b> 18  39.95

True or false?

The 1<sup>st</sup> energy level of an atom is closest to the nucleus.

The electrons of an atom are located in regions (Cloud) around the nucleus called \_\_\_\_\_.

Electrons around the nucleus are called energy levels. The 1<sup>st</sup> energy level

- a) Is furthest from the nucleus
- b) Is closest to the nucleus
- c) Holds the most electrons
- d) Needs more than two electrons to fill it up

The atoms in a **column** of the periodic table have

- a) The same abbreviation
- b) The same number of energy levels
- c) The same number of electrons
- d) The same number of electrons in the outer energy level