

HW 14

Balance the following equations.

Remember that the number of atoms of each type should be the same on the left and the right parts of the equation!

E.g. **unbalanced equation**: $\text{K} + \text{H}_2\text{O} \rightarrow \text{KOH} + \text{H}_2$

(there are 2 H on the left but 3 on the right. There should be as many atoms of each element after the reaction as before it – remember mass conservation law!)

The same equation balanced: $2\text{K} + 2\text{H}_2\text{O} \rightarrow 2\text{KOH} + \text{H}_2$ (there are 2 K on each side, 2 O on each side and 4 H on each side).

